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THE FIRST EXAMPLE OF A GALLIUM-ANTIMONY MIXED-BRIDGE COMPOUND**

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Abstract- The 1:1 mole ratio reaction of *t*-Bu₃Ga with Sb(SiMe₃)₃ in hexane yields the expected Lewis acid-base adduct *t*-Bu₃Ga•Sb(SiMe₃)₃ (**1**). The 1:1 mole ratio dehalosilylation reaction of *t*-Bu₂GaCl with Sb(SiMe₃)₃ yields the dimeric compound [*t*-Bu₂GaSb(SiMe₃)₂]₂ (**2**). The mixed-bridge compound *t*-Bu₂ $\overline{\text{GaSb(SiMe}_3)_2\text{Ga}}$ (*t*-Bu)₂Cl (**3**) was isolated from both the 2:1 reaction of *t*-Bu₂GaCl with Sb(SiMe₃)₃ and the 2:1 equilibration of *t*-Bu₂GaCl with **2**. These new gallium-antimony compounds have been characterized through multinuclear solution NMR spectroscopy, partial elemental analysis, and single-crystal X-ray structural analysis. In addition, compound **2** was found to produce nanocrystalline GaSb with an approximate average particle size of 9 nm upon thermolysis at 400 °C under vacuum.

Introduction

The bulk of our recent research into potential precursor compounds to 13-15 materials has focused primarily on Group 13 metals and the pnictogens phosphorus and arsenic¹. These compounds have been isolated largely in the form of 1:1 Lewis acid-base adducts² and dimeric compounds containing an $\overline{\text{M-E-M-E}}$ type core^{1b,3}. These dimeric compounds are generally obtained through either dehalosilylation or lithium halide elimination reactions, by combining R₂MX (R = alkyl, aryl; X = Cl, Br) in a 1:1 ratio with E(SiMe₃)₃ or LiE(SiMe₃)₂ (E = P, As), respectively. Less common are compounds containing a mixed-bridge type core ring of the form $\overline{\text{M-E-M-X}}$ (X = Cl^{3d,4} or Br^{4b}), obtained from dehalosilylation of R₂MX and E(SiMe₃)₃ in a 2:1 ratio, or equilibration of a dimeric compound with two equivalents of R₂MX.

In an attempt to extend our methodologies to the formation of precursors to Group 13-antimonides, we have recently reported the synthesis and characterization of the adducts $\text{Et}_3\text{Ga}\bullet\text{Sb}(\text{SiMe}_3)_3$ and $(\text{Me}_3\text{SiCH}_2)_3\text{In}\bullet\text{Sb}(\text{SiMe}_3)_3$, as well as the oligomeric compounds $[(\text{Me}_3\text{CCH}_2)_2\text{GaSb}(\text{SiMe}_3)_2]_n$ and $[(\text{Me}_3\text{SiCH}_2)_2\text{InSb}(\text{SiMe}_3)_2]_2$.⁵ We, as well as others, have also demonstrated that nanocrystalline GaSb is obtained from the 1:1 dehalosilylation reaction of GaCl_3 and $\text{Sb}(\text{SiMe}_3)_3$ in pentane solution⁶. These results show that further study into this area is indeed warranted, as does the lack of known Group 13-antimony compounds⁷⁻¹⁰. In an effort to further expand the library of known Group 13-antimony compounds, herein we report the synthesis and complete characterization of $t\text{-Bu}_3\text{Ga}\bullet\text{Sb}(\text{SiMe}_3)_3$ (**1**), $[t\text{-Bu}_2\text{GaSb}(\text{SiMe}_3)_2]_2$ (**2**), and $t\text{-Bu}_2\overline{\text{GaSb}(\text{SiMe}_3)_2\text{Ga}(t\text{-Bu})_2}\text{Cl}$ (**3**), the first gallium-antimony mixed-bridge compound. In addition, the thermal decomposition of **2** was examined in order to assess the suitability of these compounds as precursors to GaSb.

Experimental Section

General Considerations: All manipulations of air- and moisture-sensitive materials were performed in a Vacuum Atmospheres HE-493 Dri-Lab containing an argon atmosphere or by standard Schlenk techniques. Hexane and toluene were distilled over sodium/potassium alloy under dry dinitrogen. $t\text{-Bu}_3\text{Ga}$ ¹¹, $t\text{-Bu}_2\text{GaCl}$ ¹¹, and $\text{Sb}(\text{SiMe}_3)_3$ ¹² were prepared from literature procedures. GaCl_3 was purchased from Strem Chemicals and used as received. The integrity of all starting materials was confirmed using ^1H NMR spectroscopy. ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR spectra were recorded on a Varian Unity XL-400 spectrometer operating at 400 and 100.6 Mhz, respectively. ^1H and $^{13}\text{C}\{^1\text{H}\}$ spectra were

referenced to TMS using the residual protons or carbons of deuterated benzene at δ 7.15 or δ 128.0, respectively. All NMR samples were prepared in 5-mm tubes which were septum-sealed under argon. Melting points (uncorrected) were obtained with a Thomas-Hoover Uni-melt apparatus, using capillaries that were flame-sealed under argon. Elemental analyses (EA) were performed by E+R Microanalytical Laboratory, Inc., Corona, NY. Powder X-ray diffraction (XRD) data were collected on a Phillips XRG-3000 diffractometer utilizing Cu-K α radiation. IR spectra of volatile gases were acquired using a gas cell on a BOMEM Michelson MB-100 FT-IR spectrometer. TGA/DTA analyses were obtained on a TA Instruments SDT 2960 simultaneous TGA/DTA apparatus.

Preparation of *t*-Bu₃Ga•Sb(SiMe₃)₃ (1): *t*-Bu₃Ga (0.241 g, 1.00 mmol) in 20 mL of hexane was added to a 250 mL Schlenk flask equipped with a stir bar and Teflon valve. Sb(SiMe₃)₃ (0.341 g, 1.00 mmol) in 20 mL of pentane was added slowly via pipet and the clear, slightly red solution was stirred overnight at room temperature. The solvent was reduced *in vacuo* to yield a red liquid, which was cooled to -30 °C to produce brown prismatic crystals of **1**, suitable for X-ray analysis (0.533 g, 92% yield). m.p. 103 - 107 °C. Anal. Calcd. (found) for C₂₁H₅₄GaSbSi₃: C, 43.31 (43.13); H, 9.35 (9.27). ¹H NMR: δ 0.41 (s, 27H, -SiMe₃), δ 1.28 (s, 27H, -CH₃). ¹³C NMR{¹H}: δ 4.86 (s, -SiMe₃), δ 32.11 (s, -CH₃).

Preparation of [*t*-Bu₂GaSb(SiMe₃)₂]₂ (2): *t*-Bu₂GaCl (0.242 g, 1.10 mmol) dissolved in 25 mL of hexane was added to a 250 mL Schlenk flask equipped with a stir-bar and Teflon valve. Sb(SiMe₃)₃ (0.375 g, 1.10 mmol) dissolved in 25 mL of hexane was

added to the flask dropwise via pipet, resulting in a clear, light yellow solution which was stirred for 1 d at room temperature. After this period, the solution had taken on a dark yellow color. The solution volume was reduced *in vacuo*, and then stored at -30 °C for several days. Colorless crystals of **2**, suitable for X-ray analysis were isolated (0.342 g, 45% yield). m.p. 220 °C (dec. to a red liquid). Anal. Calc. (found) for $C_{28}H_{72}Ga_2Sb_2Si_4$: C, 37.20 (37.10); H, 8.03 (7.95). 1H NMR: δ 0.37 (s, 36H, -SiMe₃), δ 1.21 (s, 36H, CH₃-). $^{13}C\{^1H\}$ NMR: δ 4.79 (s, -SiMe₃), δ 30.77 (s, -CH₃).

Preparation of $t\text{-Bu}_2\text{GaSb}(\text{SiMe}_3)_2\text{Ga}(t\text{-Bu})_2\text{Cl}$ (3**):**

(Method 1): [$t\text{-Bu}_2\text{GaSb}(\text{SiMe}_3)_2$]₂ (0.174 g, 0.20 mmol) partially dissolved in 20 mL of hexane was added to a 250 mL Schlenk flask equipped with a stir-bar and Teflon valve. A small amount of toluene (approximately 5 mL) was added to increase the solubility. $t\text{-Bu}_2\text{GaCl}$ (0.084 g, 0.4 mmol) in 20 mL of toluene was added via pipet, resulting in a clear amber solution which was allowed to stir at room temperature for 2.5 d. The solvent was reduced *in vacuo* yielding a red liquid which was allowed to evaporate at -30 °C for 3 d, producing colorless, prismatic crystals of **3**, suitable for X-ray analysis (.097 g, 36% yield). m.p. 174 - 182 °C (dec. to a red liquid). Anal. Calc. (found) for $C_{22}H_{54}Ga_2SbSi_2Cl$: C, 39.35 (39.12); H, 8.11 (7.98). 1H NMR: δ 0.49 (s, 18H, -SiMe₃), δ 1.32 (s, 36H, CH₃-). $^{13}C\{^1H\}$ NMR: δ 6.05 (s, -SiMe₃), δ 30.82 (s, -CH₃).

(Method 2): $t\text{-Bu}_2\text{GaCl}$ (0.316 g, 1.44 mmol) dissolved in 25 mL of hexane was added to a 250 mL Schlenk flask. $Sb(\text{SiMe}_3)_3$ (0.245 g, 0.72 mmol) dissolved in 25 mL of hexane was added to the flask dropwise via pipet, resulting in a clear, light yellow solution which was stirred for 5 d at room temperature. After 1 d, the solution began to take on a

golden color, and gradually continued to darken. The solution volume was reduced *in vacuo*, and then evaporated at $-30\text{ }^{\circ}\text{C}$. Colorless crystals of **3** formed, as confirmed by ^1H NMR analysis (0.099 g, 21% yield).

Thermal Decomposition of 2. (a) TGA Analysis: Figure 4 shows the weight loss of **2** under nitrogen flow with a $5\text{ }^{\circ}\text{C}/\text{min}$ heating rate. Segment 1: $150 - 200\text{ }^{\circ}\text{C}$ with 44 % weight loss. Segment 2: $250 - 400\text{ }^{\circ}\text{C}$ with 5 % weight loss. Total observed weight loss: 52 %. Total calculated weight loss for GaSb formation: 57.6 %

(b) Pyrolysis at $400\text{ }^{\circ}\text{C}$: The sample (0.227 g, 0.50 mmol based on the monomeric unit) was loaded into a sublimator and heated under dynamic vacuum as follows: $175\text{ }^{\circ}\text{C}$, 15 min; $200\text{ }^{\circ}\text{C}$, 10 min; $400\text{ }^{\circ}\text{C}$, 12 h. A black residue was present on the cold finger, and a metallic mirror had formed on the base of the sublimator wall during this time. A gray/black powder was recovered (0.061g, 64 % yield based on GaSb), and its identity confirmed through comparison of the d-spacings and line intensities obtained by XRD analysis with those of GaSb (JCPDS file 7-215). The approximate average particle size of 9 nm was calculated using the Scherrer equation. Anal. Calc. (found) for GaSb: Ga, 36.41 (35.17); Sb, 63.59 (61.96). The Ga:Sb ratio was 1.00:1.01. A second sample of **2** (0.198 g, 0.44 mmol based on the monomeric unit) was decomposed as described above, with the exception that static vacuum was used and a $-196\text{ }^{\circ}\text{C}$ cold trap was attached in an attempt to study the volatiles formed during decomposition. The results were as follows: noncondensable CH_4 and H_2 , approximately 0.5 mmol total (0.3 mmol CH_4 and 0.2 mmol H_2); condensable HSiMe_3 and iso-butylene, approximately 0.53 mmol total. CH_4 , HSiMe_3 , and iso-butylene were identified by IR spectroscopy. The

0.023 g of black powder recovered (27 % yield based on GaSb) was identified through XRD and analyzed for impurities. Anal. Calc. (found) for this second sample: C, 0.0 (2.28); H, 0.0 (0.99).

X-ray structural solution and refinement

Crystal, data collection, and refinement parameters are given in Table 1, while selected bond lengths and angles are presented in Figures 1, 2, and 3. ORTEP diagrams showing the solid state conformations of **1** - **3** can be found in Figures 1 through 3, respectively. Compound **1**: A suitable crystal was selected and mounted in a nitrogen-flushed glass capillary. The unit-cell parameters were obtained by the least-squares refinement of the angular settings of 24 reflections ($20^\circ \leq 2\theta \leq 25^\circ$). The systematic absences in the diffraction data are consistent for the space groups $Pna2_1$ or $Pnma$. The E -statistics and the presence of a molecular mirror plane in the molecule suggested the centrosymmetric option which was subsequently verified by chemically reasonable and computationally stable results of refinement. The structure was solved using direct methods, completed by subsequent difference Fourier syntheses and refined by full-matrix least-squares procedures. Absorption corrections were not necessary because there was less than 10% variation in the integrated ψ -scan intensities. The molecule is located on a mirror plane. All non-hydrogen atoms were refined with anisotropic displacement coefficients and hydrogen atoms were treated as idealized contributions. All software and sources of the scattering factors are contained in the SHELXTL (Version 5.3) program library (G. Sheldrick, Siemens XRD, Madison, WI).

Compounds **2** and **3**: Single crystals of **2** and **3** were mounted on a glass fiber with

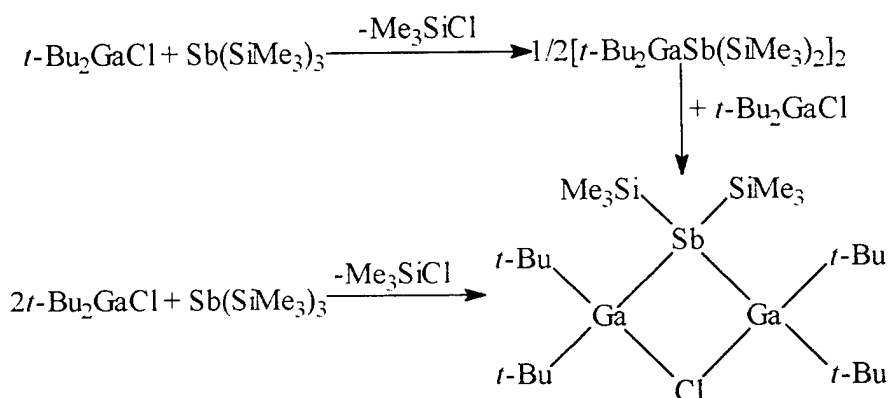
a viscous oil under a stream of cold dinitrogen. X-ray intensity data were recorded at -135 °C on a Siemens SMART CCD diffractometer utilizing graphite-monochromated Mo-K α radiation ($\lambda = 0.71073$ Å) and the structures were solved by direct methods. Full-matrix least-squares refinement with weights based upon counting-statistics was performed. Hydrogen atoms were incorporated at their calculated positions using a riding model in the later iterations of refinement which converged at $R = 0.032$ ($R_w = 0.041$) for **2** and $R = 0.050$ ($R_w = 0.053$) for **3**. A final difference Fourier synthesis revealed no unusual features. Crystallographic calculations were performed using the NRCVAX¹³ suite of structure determination programs. For all structure-factor calculations, neutral atom scattering factors and their anomalous dispersion corrections were taken from ref. 14.

Results and Discussion

The 1:1 mole ratio reaction of *t*-Bu₃Ga and Sb(SiMe₃)₃ affords in high yield the expected Lewis acid-base adduct *t*-Bu₃Ga•Sb(SiMe₃)₃ (**1**). The 1:1 mole ratio dehalosilylation reaction of *t*-Bu₂GaCl with Sb(SiMe₃)₃ yielded the dimeric compound [*t*-Bu₂GaSb(SiMe₃)₂]₂ (**2**), while the 2:1 mole ratio reaction of *t*-Bu₂GaCl with Sb(SiMe₃)₃ resulted in the formation of *t*-Bu₂Ga $\overline{\text{Sb(SiMe}_3)_2\text{Ga(t-Bu)}_2\text{Cl}}$ (**3**) after 5 days. Initially, the reaction was stopped after 1 day, and the only products isolated were **2** and *t*-Bu₂GaCl. Interestingly, the attempted formation of $\overline{\text{Ga-E-Ga-Cl}}$ core structures through 2:1 reactions yields a wide variety of results. The 2:1 reaction of (Me₃CCH₂)₂GaCl with P(SiMe₃)₃ fails to yield the mixed-bridge, instead forming [(Me₃CCH₂)(Cl)GaP(SiMe₃)₂]₂^{3d}. In the case of (Me₃CCH₂)₂GaCl reacted 2:1 with

$\text{As}(\text{SiMe}_3)_3$, the expected mixed-bridge was formed after 5 days^{4c}, comparable to the time involved in the formation of **3**. Also noteworthy is the reaction of $t\text{-Bu}_2\text{GaCl}$ with $\text{As}(\text{SiMe}_3)_3$, which failed to undergo dehalosilylation in 1:1 or 2:1 ratios, and the mixed-bridge was obtained only from reaction of $[t\text{-Bu}_2\text{GaAs}(\text{SiMe}_3)_2]_2$ with $t\text{-Bu}_2\text{GaCl}$ ¹⁵. The mixed-bridge **3** was also obtained from the equilibration reaction of $t\text{-Bu}_2\text{GaCl}$ with **2** in a 2:1 mole ratio. These reactions are summarized in Scheme 1. Compound **3** is interesting due to the possibility of metathetical reactions involving the ring chlorine atom and a $\text{LiE}(\text{SiMe}_3)_2$ ($\text{E} = \text{P}, \text{As}$) salt to form the rare mixed-pnictogen structure. This possibility is currently under investigation. The solid-state structures of **1**, **2**, and **3** have been confirmed through single crystal X-ray analysis.

Scheme 1



Compound **1** crystallizes in the orthorhombic space group $Pnma$. The Ga-Sb bond length of 3.027(2) Å in this structure is significantly longer than the bond length of 2.846(5) Å reported for the adduct $\text{Et}_3\text{Ga} \cdot \text{Sb}(\text{SiMe}_3)_3$ ⁵. This discrepancy can be explained by the extreme steric bulk of the t -butyl- group in comparison to the ethyl- group, which prevents close approach of the $\text{Sb}(\text{SiMe}_3)_3$ moiety. It should also be noted

that the ligands bound to the metal centers adopt a staggered conformation in relation to one another.

Crystals of the dimeric compound **2** belong to the monoclinic space group $C2/c$ and lie on a two-fold axis. The average Ga-Sb bond length of 2.7666(7) Å found in **2** is slightly longer than the average Ga-Sb bond length of 2.66 Å reported for the ring compound $[\text{Cl}_2\text{GaSb}(\text{t-Bu})_2]_3$ ⁷. Again, this is probably due to the steric bulk exerted by the *t*-butyl- groups. The average endocyclic bond angles in **2** are 94.45(3)° for Ga-Sb-Ga and 85.549(24)° for Sb-Ga-Sb. The gallium and antimony atoms in this compound reside in distorted tetrahedral environments and the core ring is planar.

The mixed-bridge compound **3** forms crystals belonging to the triclinic space group $P\bar{1}$. The average Ga-Sb bond length of 2.7336(11) Å compares well with the analogous bond length in **2** (*vide supra*). The core ring of this compound is again planar, with average endocyclic Ga-Sb-Ga, Ga-Cl-Ga, and Sb-Ga-Cl ring angles of 85.68(3)°, 98.18(9)°, and 88.07(6)°, respectively. The gallium and antimony atoms again reside in distorted tetrahedral environments, with the influence of the *t*-butyl- groups demonstrated in the appropriate bond angles.

In an attempt to assess the utility of these compounds as precursors to GaSb, the thermal decomposition behavior of **2** was closely examined. This compound was found to produce reasonably pure GaSb in moderate to low yield. In addition, analysis of the volatile products produced during decomposition suggests that the formation of GaSb occurs through a β -hydride elimination pathway. This is evidenced through the condensation of HSiMe_3 and iso-butylene in the cold trap, the expected products of such

an elimination process. Both compounds were identified by IR spectroscopy. If this elimination occurred cleanly, the theoretical amount of HSiMe_3 and iso-butylene collected would be 1.76 mmol total (based on 0.44 mmol of starting material). The fact that approximately a third of this amount, 0.53 mmol, was observed suggests that the elimination was not complete, and a secondary process was also involved. Two other noncondensable gasses were produced during the pyrolysis, namely CH_4 and H_2 , presumably from decomposition of the $-\text{SiMe}_3$ groups. It is possible to envision a two step mechanism for **2** undergoing β -hydride elimination and subsequently forming GaSb, with elimination of iso-butylene as step 1 followed by elimination of trimethylsilane as step 2. However, the TGA plot for this process, shown in Figure 4, indicates only one major elimination occurring, in the range of 150 - 200 °C. This suggests that steps 1 and 2 occur simultaneously. The small (approximately 10 %), broad weight loss observed from 250 - 400 °C is probably due to evolution of residual silicon and carbon containing species not lost in the initial elimination, and could account for the observed CH_4 and H_2 gas. It should be noted that initial pyrolysis experiments on **1** indicate that GaSb is also formed, however the purity and yield are less than observed for **2**, and the system is still under investigation.

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Supplementary Material Available: Tables of bond distances, bond angles, anisotropic temperature factor parameters, and fractional coordinates for **1**, **2**, and **3** (26 pages). Ordering information is given on any current masthead page.

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Captions to Figures

Figure 1. Molecular structure of **1** drawn with 30% probability ellipsoids. Methyl- group carbon atoms are rendered spherically due to their high thermal activity. Bond Lengths (Å): Ga-Sb: 3.027(2); Sb-Si(1): 2.555(2); Ga-C(6): 2.007(8); Sb-Si(2): 2.566(3); Ga-C(10): 2.027(9). Bond Angles (°): Ga-Sb-Si(1): 118.53(6); Sb-Ga-C(6): 98.7(3); Ga-Sb-Si(2): 114.68(7); Sb-Ga-C(10): 102.8(3); Si(1)-Sb-Si(2): 100.65(8); C(6)-Ga-C(10): 116.7(3); Si(1)-Sb-Si(1)A: 100.75(13); C(6)-Ga-C(6)A: 117.5(5).

Figure 2. ORTEP diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme for **2**. Hydrogen atoms are omitted for clarity. Bond Lengths (Å): Ga1-Sb1: 2.7684(7); Ga1-Sb2: 2.7648(7); Ga1-C31: 2.018(6); Sb1-Si1: 2.5918(17); Ga1-C41: 2.036(5); Sb2-Si2: 2.5827(18). Bond Angles (°): Ga1-Sb1-Ga1: 94.37(3); Ga1-Sb2-Ga1: 94.53(3); Sb1-Ga1-Sb2: 85.549(24); C31-Ga1-Sb1: 115.30(18); C41-Ga1-Sb1: 109.61(16); C31-Ga1-Sb2: 110.54(17); C41-Ga1-Sb2: 112.76(18); Si1-Sb1-Ga1: 115.83(4); Si2-Sb2-Ga1: 118.74(4); C31-Ga1-C41: 118.5(3); Si1-Sb1-Si1: 96.52(6); Si2-Sb2-Si2: 96.13(6); Si1A-Sb1-Ga1: 117.98(4); Si2A-Sb2-Ga1: 115.22(5).

Figure 3. ORTEP diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme for **3**. Hydrogen atoms are omitted for clarity. Bond Lengths (Å, average): Ga1-Sb1: 2.7336(12); Ga2-Sb1: 2.7337(11); Ga1-Cl1: 2.4568(24); Ga2-Cl1: 2.462(3); Ga1-C11: 2.032(9); Ga1-C21: 2.007(9); Ga2-C31: 2.019(8); Ga2-C41: 2.007(9); Sb1-Si1: 2.599(3); Sb1-Si2: 2.599(3). Bond Angles (°, average): Ga1-Sb1-Ga2: 85.68(3); Ga1-Cl1-Ga2: 98.18(9); Sb1-Ga1-Cl1: 88.12(6); Sb1-Ga2-Cl1: 88.01(6); C11-Ga1-Sb1: 117.4(3); C11-Ga1-Cl1: 105.7(3); C21-Ga1-Sb1: 111.8(3); C21-Ga1-Cl1: 103.9(3); C11-Ga1-C21: 122.5(4); C31-Ga2-C41: 122.4(4); C31-Ga2-Sb1: 117.4(3); C31-Ga2-Cl1: 105.5(3); C41-Ga2-Sb1: 112.09(25); C41-Ga2-Cl1: 103.6(3); Si1-Sb1-Ga1: 120.81(7); Si1-Sb1-Ga2: 117.42(6); Si2-Sb1-Ga1: 117.34(7); Si2-Sb1-Ga2: 120.85(7); Si1-Sb1-Si2: 96.93(8).

Figure 4. Thermogravimetric Analysis (TGA) data for the decomposition of [*t*-Bu₂GaSb(SiMe₃)₂]₂ (**2**).

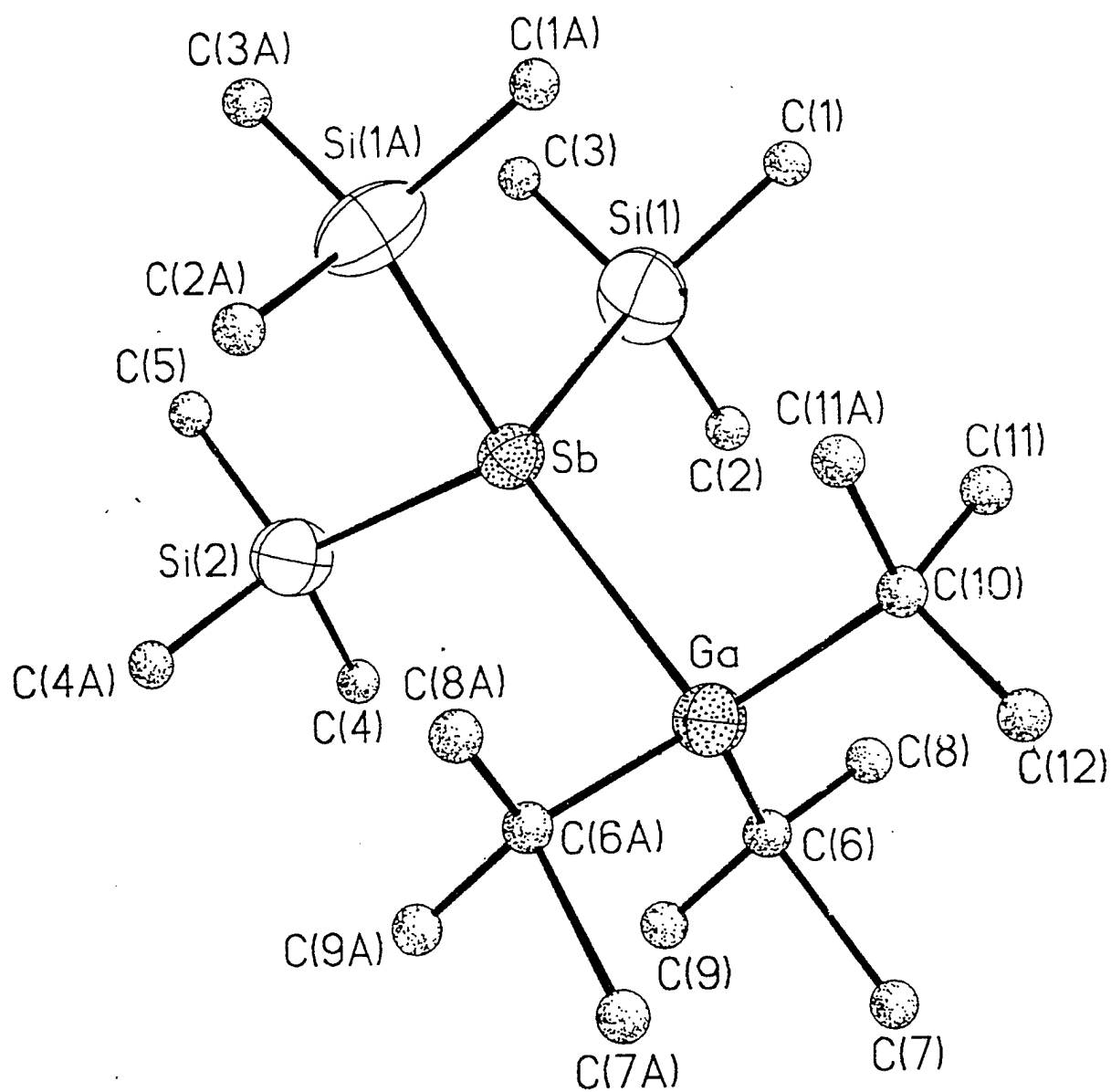


Figure 1.

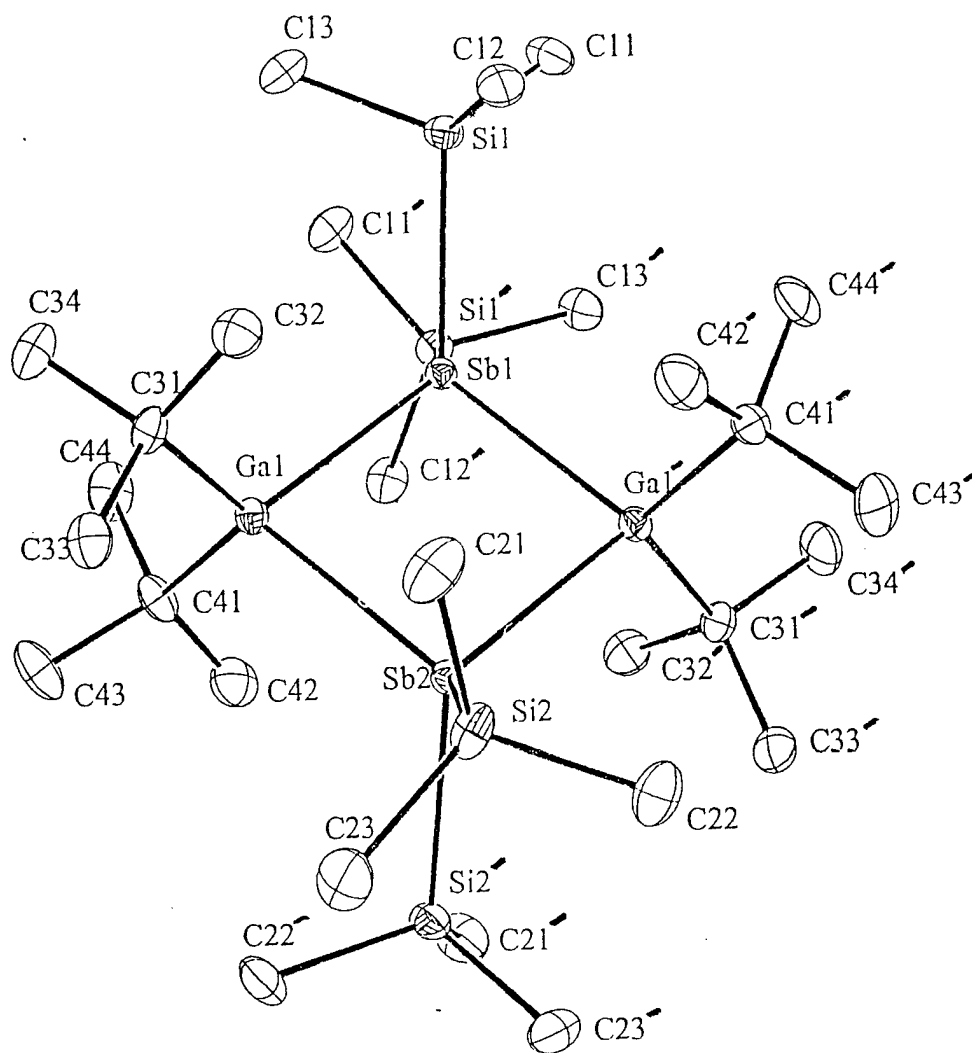


Figure 2.

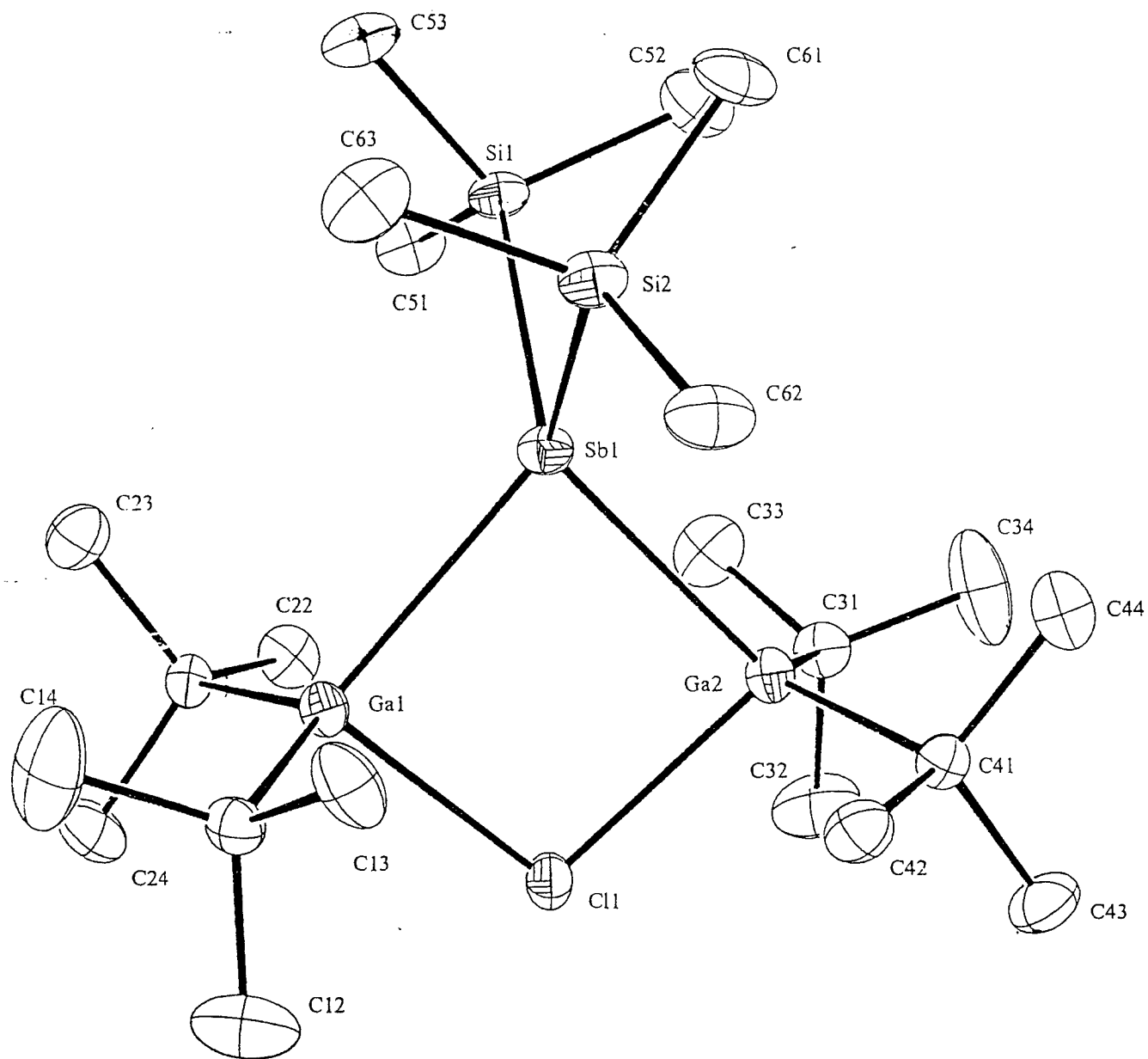


Figure 3.

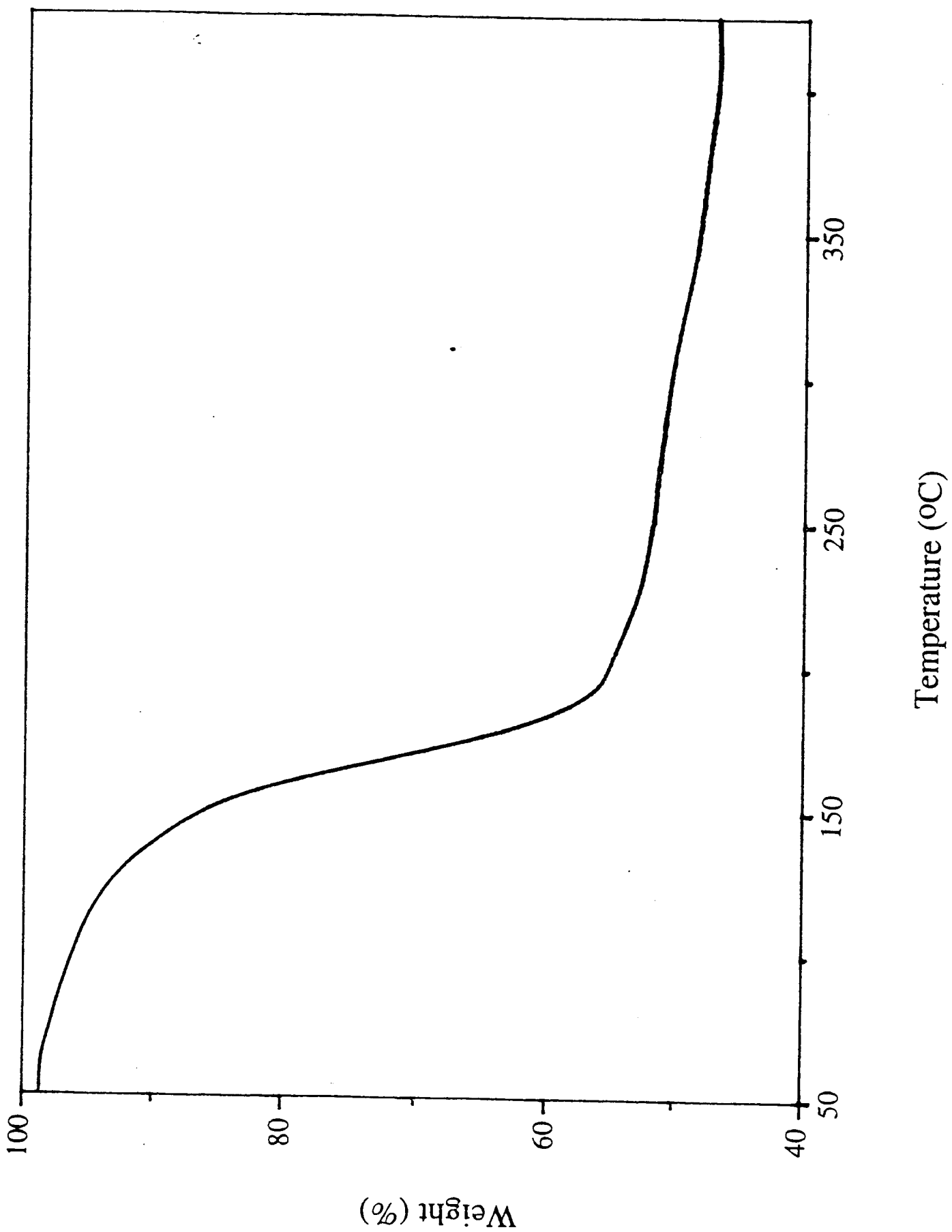


Figure 4.

Table 1. Crystallographic Data for $t\text{-Bu}_3\text{Ga}\bullet\text{Sb}(\text{SiMe}_3)_3$ (1), $[\text{t-Bu}_2\text{GaSb}(\text{SiMe}_3)_2]_2$ (2), and $t\text{-Bu}_2\text{GaSb}(\text{SiMe}_3)_2\text{Ga}(\text{t-Bu})_2\text{Cl}$ (3).

	1	2	3
formula	$\text{C}_{21}\text{H}_{54}\text{GaSbSi}_3$	$\text{C}_{28}\text{H}_{72}\text{Ga}_2\text{Sb}_2\text{Si}_4$	$\text{C}_{22}\text{H}_{54}\text{Ga}_2\text{SbSi}_2\text{Cl}$
formula weight	582.38	904.15	671.47
space group	$Pnma$	$C2/c$	$P\bar{1}$
a , Å	17.504(3)	19.614(4)	10.0358(4)
b , Å	15.923(3)	13.280(3)	11.5969(5)
c , Å	11.522(5)	18.230(3)	16.3753(7)
α , deg.	---	---	82.2250(1)
β , deg.	---	116.912(12)	72.1790(1)
γ , deg.	---	---	64.3690(1)
V , Å ³	3209(2)	4234.2(14)	1635.83(12)
Z	4	4	2
crystal color, habit	brown block	colorless block	colorless block
$D(\text{calc})$, g cm ³	1.206	1.418	1.363
$\mu(\text{MoK}\alpha)$, cm ⁻¹	17.98	26.5	26.2
temp, K	245(2)	138	138
radiation	MoK α ($\lambda = 0.71073$ Å)		
$R(F)$, %	4.31 ^a	3.2 ^b	5.0 ^b

Table 1. (cont.)

$R(wF^2)$, %	10.16 ^a	4.1 ^b	5.3 ^b
---------------	--------------------	------------------	------------------

^a Quantity minimized = $R(wF^2) = \Sigma[w(F_o^2 - F_c^2)^2] / \Sigma[wF_o^2]^2$; $R = \Sigma\Delta / \Sigma(F_o)$, $\Delta = |(F_o - F_c)|$

^b Quantity minimized = $\Sigma[w(F_o - F_c)^2] / \Sigma[wF_o^2]^2$; $R = \Sigma\Delta / \Sigma(F_o)$, $\Delta = |(F_o - F_c)|$

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